5.3 Extra Practice

How much energy is absorbed or released (identify which) when each of the following reactions takes place?

1. CH4 (g) + 2 O2 (g) 🡪CO2 (g) + 2 H2O (g)
2. 2 H2 (g) + O2 (g) 🡪2 H2O (g)
3. H2 (g) + Cl2 (g) 🡪2 HCl (g)
4. CH4 (g) + Cl2 (g) 🡪CH3Cl (g) + HCl (g)
5. N2 (g)+ 3 Cl2 (g) 🡪2 NCl3 (g)
6. Combustion of but-1-ene.
7. CH3CHCH2 + Br2 🡪 CH3CHBrCH2Br
8. Propan-2-ol 🡪 propanone + H2O